

1,4-Bis[2-(prop-1-enyl)phenoxy]butane

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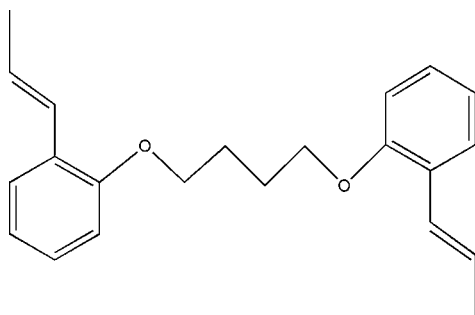
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.042; wR factor = 0.151; data-to-parameter ratio = 21.9.

The molecule of the title compound, $\text{C}_{22}\text{H}_{26}\text{O}_2$, exhibits C_i molecular symmetry with a crystallographic inversion centre at the mid-point of the central C—C bond. A kink in the molecule is defined by the torsion angle of $66.7(2)^\circ$ about this central bond of the alkyl bridge.

Related literature

For general background to the use of copolymerization reactions, see: Crivello *et al.* (1994); Roshupkin & Kurmaz (2004); Askadsky (1998).



Experimental

Crystal data

$\text{C}_{22}\text{H}_{26}\text{O}_2$
 $M_r = 322.43$
Orthorhombic, $Pbca$
 $a = 5.4501(10)$ Å
 $b = 15.825(3)$ Å
 $c = 21.889(4)$ Å
 $V = 1887.9(6)$ Å³
 $Z = 4$
Mo $K\alpha$ radiation
 $\mu = 0.07$ mm⁻¹
 $T = 296$ K
 $0.30 \times 0.20 \times 0.20$ mm

Data collection

Bruker SMART APEXII CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
 $T_{\min} = 0.979$, $T_{\max} = 0.986$
20262 measured reflections
2404 independent reflections
1427 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.040$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.042$
 $wR(F^2) = 0.151$
 $S = 1.01$
2404 reflections
110 parameters
H-atom parameters constrained
 $\Delta\rho_{\max} = 0.15$ e Å⁻³
 $\Delta\rho_{\min} = -0.11$ e Å⁻³

Data collection: *APEX2* (Bruker, 2005); cell refinement: *SAINT-Plus* (Bruker, 2001); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

We thank Professor Victor N. Khrustalev for fruitful discussions and help in this work.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: KP2326).

References

- Askadsky, A. A. (1998). *Successes Chem.* **67**, 755–787.
Bruker (2001). *SAINT-Plus* and *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
Bruker (2005). *APEX2*. Bruker AXS Inc., Madison, Wisconsin, USA.
Crivello, J. V., Carter, A. M. & Bratslavsky, S. A. (1994). *J. Polymer Sci. Part A Polymer Chem.* **32**, 2895–2909.
Roshupkin, V. P. & Kurmaz, S. V. (2004). *Successes Chem.* **73**, 247–274.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

supplementary materials

Acta Cryst. (2011). E67, o1478 [doi:10.1107/S1600536811018538]

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M. R. Bayramov, A. M. Maharramov, G. M. Mehdiyeva, S. B. Hoseinzadeh and R. K. Askerov

Comment

For giving the important operationalisation to cross-linked polymers reactions of copolymerisation of various monomers, multifunctional comonomers [1] are used. By applying this method, the hi-tech processes for the preparation of polymeric sorbents, photorezisting materials for microelectronics, the composites for the laser technics of a special purposes, fixation of metalcomplex of catalysts, *etc* were obtained (Crivello *et al.* 1994). In practice, for obtaining polymers of demanded functional properties, polymerical transformations are carried out. However, it is necessary to notice, that obtaining such cross-linked copolymers have some difficulties connected with high reactivity of cross-linking comonomers (for example, divinylbenzene), which is reflected in heterogeneity of their structure and other important physical and chemical properties. Therefore, to prepare multifunctional monomers, on the basis of alkenylphenols with two double bonds, is rather important. The molecule of the title compound, C₂₂H₂₆O₂ (I), is generated by a crystallographic inversion centre at the midpoint of the central C—C bond. A fold of the molecule is due to the twist in the central butylene bridge [O1—C10—C11—C11A torsion angle of 66.7 (2)°] (Fig. 1). Crystal packing is dominated by van der Waals interactions (Fig. 2).

Experimental

2-Propenylphenol (0.015 mol, 2 g) and KOH (0.015 mol, 0.84 g) were dissolved in 5 mL 2-propanol, then to this solution 1,4-dibromobutane (0.0043 mol, 0.93 g) was added. This mixture was stirred at 353 K within 30 min. The desired compounds (with yield of 4.7 g, 98.1%) was filtered and washed with acetone and recrystallised to obtain colourless crystals. Tmp = 353 K. The structure of the reported compound - 1,4-bis{2(1-propenyl)phenoxy}butane, also was proved by NMR-spectroscopy. FT-NMR (acetone-d₆, p.p.m.), ¹H: 1.91 d (6H,CH₃); 2.03 t (4H,CH₂); 4.1 t (4H, OCH₂); 6.15 m (2H, CH=); 6.65–7.1 m (8H, 2Ar); 7.32 d (2H,CH=). ¹³C: 18.9; 26.3; 67.6; 112.8; 121.7; 124.9; 126.1; 127.2; 127.3; 127.5; 156.4.

Refinement

The hydrogen atoms were placed in calculated positions and refined in the riding mode with fixed isotropic displacement parameters [$U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$].

Figures

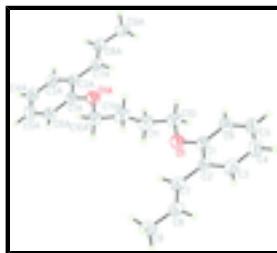


Fig. 1. The molecular structure of the title compound with the atomic numbering scheme. Displacement ellipsoids were drawn at the 30% probability level.

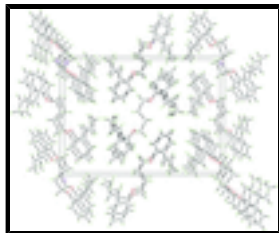


Fig. 2. Stacking of chains in the crystal packing.

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Crystal data

$C_{22}H_{26}O_2$

$M_r = 322.43$

Orthorhombic, *Pbca*

Hall symbol: -P 2ac 2ab

$a = 5.4501(10) \text{ \AA}$

$b = 15.825(3) \text{ \AA}$

$c = 21.889(4) \text{ \AA}$

$V = 1887.9(6) \text{ \AA}^3$

$Z = 4$

$F(000) = 696$

$D_x = 1.134 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 4034 reflections

$\theta = 2.6\text{--}23.1^\circ$

$\mu = 0.07 \text{ mm}^{-1}$

$T = 296 \text{ K}$

Prism, colourless

$0.30 \times 0.20 \times 0.20 \text{ mm}$

Data collection

Bruker SMART APEXII CCD
diffractometer

Radiation source: fine-focus sealed tube
graphite

φ and ω scans

Absorption correction: multi-scan
(*SADABS*; Bruker, 2001)

$T_{\min} = 0.979$, $T_{\max} = 0.986$

20262 measured reflections

2404 independent reflections

1427 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.040$

$\theta_{\max} = 28.6^\circ$, $\theta_{\min} = 1.9^\circ$

$h = -7 \rightarrow 7$

$k = -21 \rightarrow 21$

$l = -29 \rightarrow 29$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.042$

$wR(F^2) = 0.151$

$S = 1.01$

2404 reflections

110 parameters

0 restraints

Primary atom site location: structure-invariant direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.071P)^2 + 0.2P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.15 \text{ e \AA}^{-3}$

$\Delta\rho_{\min} = -0.11 \text{ e \AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.1929 (2)	0.12834 (6)	0.05285 (5)	0.0764 (3)
C1	0.1981 (3)	0.19881 (9)	0.08887 (6)	0.0641 (4)
C2	0.3717 (2)	0.19780 (9)	0.13598 (6)	0.0614 (4)
C3	0.3831 (3)	0.26839 (11)	0.17357 (8)	0.0805 (5)
H3A	0.4969	0.2696	0.2052	0.097*
C4	0.2297 (4)	0.33659 (12)	0.16510 (10)	0.0964 (6)
H4A	0.2408	0.3832	0.1909	0.116*
C5	0.0609 (4)	0.33587 (12)	0.11879 (11)	0.0958 (6)
H5A	-0.0436	0.3817	0.1134	0.115*
C6	0.0455 (3)	0.26750 (11)	0.08014 (8)	0.0826 (5)
H6A	-0.0673	0.2675	0.0483	0.099*
C7	0.5292 (3)	0.12253 (10)	0.14476 (7)	0.0655 (4)
H7A	0.4863	0.0747	0.1225	0.079*
C8	0.7203 (3)	0.11548 (10)	0.17993 (7)	0.0729 (4)
H8A	0.7671	0.1632	0.2018	0.088*
C9	0.8691 (3)	0.03876 (12)	0.18816 (7)	0.0789 (5)
H9A	0.8709	0.0234	0.2306	0.118*
H9B	1.0338	0.0494	0.1746	0.118*
H9C	0.7998	-0.0065	0.1646	0.118*
C10	0.0059 (3)	0.12144 (11)	0.00726 (8)	0.0791 (5)
H10A	0.0261	0.1653	-0.0233	0.095*
H10B	-0.1549	0.1277	0.0257	0.095*
C11	0.0297 (4)	0.03563 (11)	-0.02174 (7)	0.0827 (5)
H11A	0.1962	0.0286	-0.0366	0.099*
H11B	-0.0797	0.0325	-0.0566	0.099*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0798 (7)	0.0745 (7)	0.0750 (7)	0.0052 (5)	-0.0182 (5)	-0.0066 (5)
C1	0.0641 (8)	0.0616 (8)	0.0665 (8)	-0.0057 (6)	0.0061 (7)	0.0056 (6)
C2	0.0609 (8)	0.0622 (8)	0.0611 (8)	-0.0133 (6)	0.0091 (6)	0.0018 (6)

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C3	0.0864 (11)	0.0742 (10)	0.0807 (10)	-0.0152 (8)	0.0078 (9)	-0.0079 (8)
C4	0.1190 (15)	0.0636 (10)	0.1066 (14)	-0.0128 (10)	0.0252 (13)	-0.0162 (10)
C5	0.1026 (13)	0.0663 (11)	0.1185 (16)	0.0098 (9)	0.0153 (13)	0.0074 (10)
C6	0.0802 (10)	0.0741 (11)	0.0936 (12)	0.0050 (8)	0.0000 (9)	0.0135 (9)
C7	0.0640 (8)	0.0719 (9)	0.0605 (8)	-0.0116 (7)	-0.0001 (6)	-0.0039 (7)
C8	0.0725 (9)	0.0806 (10)	0.0657 (9)	-0.0124 (8)	-0.0047 (7)	-0.0068 (7)
C9	0.0681 (9)	0.0978 (12)	0.0708 (9)	0.0015 (8)	-0.0044 (7)	0.0010 (8)
C10	0.0803 (10)	0.0902 (11)	0.0666 (9)	-0.0064 (8)	-0.0164 (8)	0.0109 (8)
C11	0.0945 (12)	0.0979 (12)	0.0558 (8)	-0.0157 (10)	-0.0108 (8)	0.0026 (7)

Geometric parameters (\AA , $^\circ$)

O1—C1	1.3659 (17)	C7—C8	1.300 (2)
O1—C10	1.4305 (19)	C7—H7A	0.9300
C1—C6	1.382 (2)	C8—C9	1.471 (2)
C1—C2	1.399 (2)	C8—H8A	0.9300
C2—C3	1.389 (2)	C9—H9A	0.9600
C2—C7	1.481 (2)	C9—H9B	0.9600
C3—C4	1.378 (3)	C9—H9C	0.9600
C3—H3A	0.9300	C10—C11	1.505 (2)
C4—C5	1.369 (3)	C10—H10A	0.9700
C4—H4A	0.9300	C10—H10B	0.9700
C5—C6	1.376 (3)	C11—C11 ⁱ	1.511 (3)
C5—H5A	0.9300	C11—H11A	0.9700
C6—H6A	0.9300	C11—H11B	0.9700
C1—O1—C10	118.67 (12)	C7—C8—C9	125.81 (15)
O1—C1—C6	123.36 (14)	C7—C8—H8A	117.1
O1—C1—C2	115.47 (12)	C9—C8—H8A	117.1
C6—C1—C2	121.17 (15)	C8—C9—H9A	109.5
C3—C2—C1	117.23 (14)	C8—C9—H9B	109.5
C3—C2—C7	122.96 (14)	H9A—C9—H9B	109.5
C1—C2—C7	119.78 (12)	C8—C9—H9C	109.5
C4—C3—C2	121.55 (18)	H9A—C9—H9C	109.5
C4—C3—H3A	119.2	H9B—C9—H9C	109.5
C2—C3—H3A	119.2	O1—C10—C11	107.57 (13)
C5—C4—C3	120.05 (17)	O1—C10—H10A	110.2
C5—C4—H4A	120.0	C11—C10—H10A	110.2
C3—C4—H4A	120.0	O1—C10—H10B	110.2
C4—C5—C6	120.20 (18)	C11—C10—H10B	110.2
C4—C5—H5A	119.9	H10A—C10—H10B	108.5
C6—C5—H5A	119.9	C10—C11—C11 ⁱ	112.91 (17)
C5—C6—C1	119.78 (18)	C10—C11—H11A	109.0
C5—C6—H6A	120.1	C11 ⁱ —C11—H11A	109.0
C1—C6—H6A	120.1	C10—C11—H11B	109.0
C8—C7—C2	127.62 (14)	C11 ⁱ —C11—H11B	109.0
C8—C7—H7A	116.2	H11A—C11—H11B	107.8
C2—C7—H7A	116.2		
C10—O1—C1—C6	-5.9 (2)	C3—C4—C5—C6	-0.7 (3)

C10—O1—C1—C2	174.45 (13)	C4—C5—C6—C1	1.1 (3)
O1—C1—C2—C3	179.89 (13)	O1—C1—C6—C5	179.48 (15)
C6—C1—C2—C3	0.3 (2)	C2—C1—C6—C5	-0.9 (2)
O1—C1—C2—C7	-1.52 (18)	C3—C2—C7—C8	-12.1 (2)
C6—C1—C2—C7	178.86 (14)	C1—C2—C7—C8	169.40 (15)
C1—C2—C3—C4	0.2 (2)	C2—C7—C8—C9	178.78 (15)
C7—C2—C3—C4	-178.36 (15)	C1—O1—C10—C11	-175.35 (13)
C2—C3—C4—C5	0.0 (3)	O1—C10—C11—C11 ⁱ	66.7 (2)

Symmetry codes: (i) $-x, -y, -z$.

Fig. 1

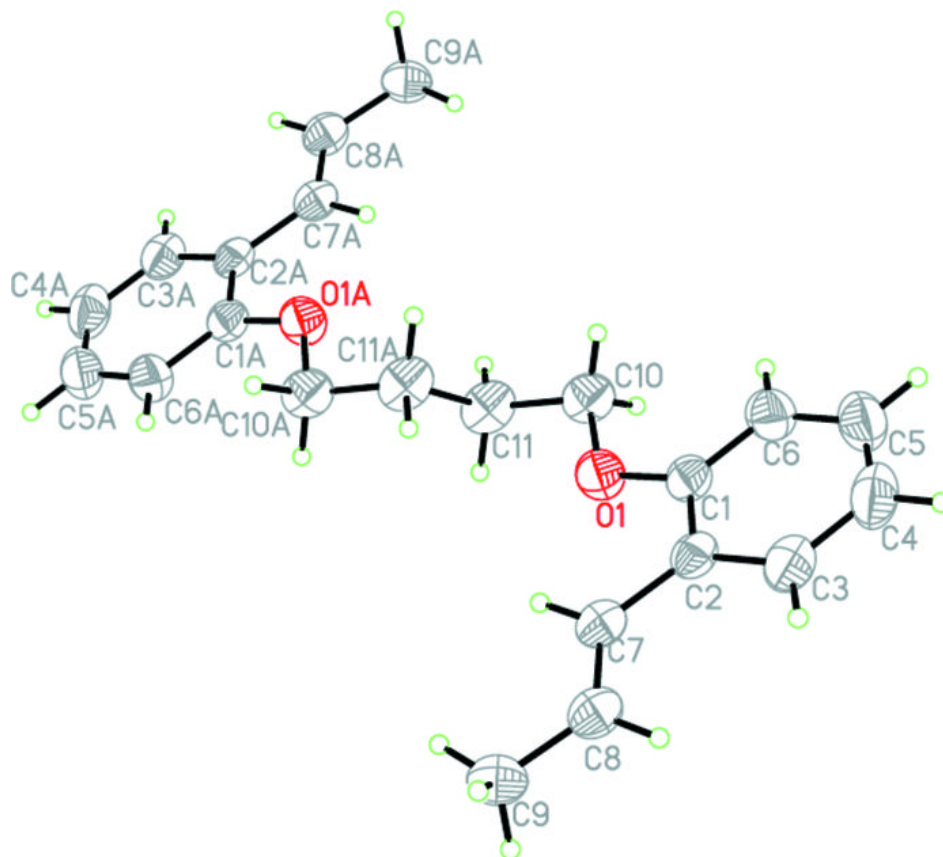


Fig. 2

